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of hydrogen-1 are in a 1.007-g sample? Recall that 1.007 amu is  
the mass of one atom of hydrogen-1. Round your answer to two  
significant digits. 1.007 g 1 H-1 atom  $\approx$  1.672  $\times$  10<sup>24</sup> g a. 6.023  
10<sup>23</sup> H-1 atoms 4.

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b. calorie e. specific heat c. joule f. heat \_\_\_\_ 1. quantity of heat needed to raise the temperature of 1 g of water by 1 C \_\_\_\_ 2. SI unit of energy \_\_\_\_ 3. energy transferred between 2 objects because of temperature difference

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