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Section 18.1 Assessment page 643 5. Explain why many Lewis acids and bases are not classified as Arrhenius or Brønsted-Lowry acids and bases. A Lewis acid is an electron pair acceptor and a Lewis base is an electron pair donor. A Lewis acid may not have an ionizable hydrogen ion or hydroxide

ion to qualify as an Arrhenius acid or base.

Acids and Bases

Accounting Chapter 11: Yes: 03/29/19: 82 students verified as accurate: Accounting Chapter 2: Yes: 04/09/19: 24 students verified as accurate: Accounting Chapter 3: Yes: 04/30/19: 89 students verified as accurate: Accounting Chapter 4: Yes: 05/05/19: 19 students verified as accurate: Accounting Chapter 5 Homework: Yes: 05/18/19: 3 students ...

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